

SCIENCE QUIZ MAY 2022

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1. The removal of oxygen from a substance is called:
(A) oxidation (B) corrosion (C) reduction (D) rancidity
2. In the context of redox reaction, the removal of hydrogen from a substance is known as:
(A) oxidation (B) dehydration (C) reduction (D) dehydrogenation
3. The chemical reaction involved in the corrosion of iron metal is that of:
(A) oxidation as well as displacement (B) reduction as well as combination
(C) oxidation as well as communication (D) reduction as well as the displacement
4. The term used to indicate the development of unpleasant smell and taste in fat and oil containing foods due to aerial oxidation is:
(A) acidity (B) radioactivity (C) rabidity (D) rancidity
5. In order of prevent the spoilage of potato chips, they are packed in plastic bags in an atmosphere of:
(A) chlorine (B) hydrogen (C) nitrogen (D) oxygen
6. A white precipitate can be obtained by adding dilute sulphuric acid to:
(A) copper sulphate solution (B) sodium chloride solution
(C) barium chloride solution (D) sodium sulphate solution
7. A white precipitate will be formed if we add common salt solution to:
(A) barium nitrate solution (B) potassium nitrate solution
(C) silver nitrate solution (D) magnesium nitrate solution
8. Consider the following equation of the chemical reaction of a metal M: $4M + 3O_2 \rightarrow 2M_2O_3$ This equation represents:
(A) Combination reaction as well as reduction reaction
(B) decomposition reaction as well as oxidation reaction
(C) oxidation reaction as well as displacement reaction
(D) combination reaction as well as oxidation reaction
9. The process of respiration is:
(A) an oxidation reaction which is endothermic
(B) a reduction reaction which is exothermic
(C) a combination reaction which is endothermic
(D) an oxidation reaction which is exothermic

10. Which of the following can be decomposed by the action of light?
(A) sodium chloride (B) potassium chloride (C) silver chloride (D) cuprous chloride
11. Consider the reaction: potassium bromide (aq) + Silver Nitrate (aq) \rightarrow potassium nitrate + silver bromide. This is an example of:
(A) decomposition reaction (B) combination reaction
(C) double displacement reaction (D) displacement reaction
12. You are given the following chemical equation: $\text{Mg (s)} + \text{CuO (s)} \rightarrow \text{MgO (s)} + \text{Cu (s)}$.
This equation represents:
(A) decomposition reaction as well as displacement reaction
(B) combination reaction as well as double displacement reaction
(C) redox reaction as well as displacement reaction
(D) double displacement reaction as well as Redox reaction
13. One of the following does not happen during a chemical reaction. This is:
(A) breaking of old chemical bonds and formation of new chemical bond
(B) formation of new substances with entirely different properties
(C) atoms of one element change into those of another element to form new products
(D) rearrangement of atoms takes place to form new products
14. Which of the following does not involve a chemical reaction?
(A) digestion of food in our body (B) process of respiration
(C) burning of candle wax when heated (D) melting of candle wax on heating
15. You are given the solution of lead nitrate. In order to obtain a yellow precipitate you should mix with it a solution of:
(A) potassium chloride (B) potassium nitride
(C) potassium sulphide (D) potassium iodide
16. An acid which can decolorize purple colour potassium permanganate solution is:
(A) sulfuric acid (B) citric acid (C) carbonic acid (D) hydrochloric acid
17. The chemical reaction between two substances is characterized by a change in colour from orange to green. These two substances are most likely to be:
(A) potassium dichromate solution and Sulphur Dioxide
(B) potassium permanganate solution and Sulphur Dioxide
(C) potassium permanganate solution and lemon juice
(D) potassium dichromate solution and carbon dioxide
18. The chemical reaction between quicklime and water is characterized by:
(A) evolution of hydrogen gas (B) formation of slaked lime precipitate
(C) change in temperature of mixture (D) change in colour of the product.

19. Out of the following is an endothermic reaction. This is:
 (A) combination of carbon and oxygen to form carbon monoxide
 (B) combination of Nitrogen and oxygen to form nitrogen monoxide
 (C) combination of glucose and oxygen to form carbon dioxide and water
 (D) combination of zinc and hydrochloric acid to form zinc chloride and hydrogen
20. Which of the following is not an endothermic reaction?
 (A) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ (B) $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 (C) $6\text{CO}_2 + 6\text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2$ (D) $\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$
21. One of the following is an exothermic reaction. This is:
 (A) electrolysis of water (B) conversion of limestone in to quicklime
 (C) process of respiration (D) process of photosynthesis
22. The chemical equation balance to satisfy one of the following the laws in chemical reactions. This law is known as:
 (A) law of conservation of momentum (B) law of conservation of mass
 (C) law of conservation of motion (D) law of conservation of magnetism

ANSWERS

21. (C) 22. (B)
11. (C) 12. (C) 13. (C) 14. (D) 15. (D) 16. (B) 17. (A) 18. (C) 19. (B) 20. (D)
1. (C) 2. (A) 3. (C) 4. (D) 5. (C) 6. (C) 7. (D) 8. (D) 9. (D) 10. (C)